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Chemical equilibrium is established when the forward rate of reaction is equal to the reverse rate of reaction. a. TRUE b. FALSE. View Answer. For a reaction $N_2 + O_2 \rightarrow 2NO$. If $K_c = 10^{-30}$...

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Chemical equilibrium (definition) The state reached when the concentration of reactants and products remain constant over time What is the general, reversible reaction for equilibrium?

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The law of chemical equilibrium states that at a given pressure, a chemical system may reach a state in which a particular ratio of reactant to product concentrations has a constant value. 7. The equation $H_2(g) + I_2(g) \rightleftharpoons 2HI(g)$ is an example of a homogeneous equilibrium. 8.

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In the following reaction list the equilibrium shift (right or left) $PCl_5 + \text{heat} \rightleftharpoons PCl_3 + Cl_2$, The addition of PCl_3 . Increasing the temperature 3. Decreasing Cl_2 4. Increasing the pressure 5. Decreasing the temperature Know the factors that can affect the rate of a reaction rcvseve.

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system in equilibrium . Find the reaction . quotient, Q, (K_{eq}) for a reaction equation . Solve for Q given . initial concentrations or pressures. Using . LeChatlier's principle, determine the direction of shift in an equilibrium when conditions are changed (concentration, pressure, phase change) Setup a . K_{sp} expression for a slightly soluble. product. Calculate the

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This model should show the three states that occur in the chemical equilibrium process. As students work, walk around to offer suggestions and guide. Have students:

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***If concentration of a substance is added the reaction will shift to reestablish equilibrium by consuming part of the added substance. ***If concentration of a substance is removed the equilibrium will shift in the direction to produce more of that substance. ***In simple terms, When you remove something, it will shift in the direction moved.

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248 Study Guide for An Introduction to Chemistry Exercise 16.2 – Equilibrium Constant Calculation: Ethanol, C₂H₅OH, can be made from the reaction of ethylene gas, C₂H₄, and water vapor.A mixture of C₂H₄(g) and H₂O(g) is allowed to come to equilibrium in a container at 110 C, and the partial pressures of the gases are found

Chapter 16 - The Process of Chemical Reactions

Chemical equilibrium is the point at which the rate of the forward reaction equals the rate of the reverse reaction. Does a chemical reaction actually stop once equilibrium is reached? Chemical reactions never stop, even when equilibrium is reached.

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